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Questions:

2 Marks

1. How many moles of atoms are present in 9.0 g of aluminum?
2. An organometallic compound on analysis was found to contain, C = 64.4%, H = 5.5 %, and Fe = 29.9%. Determine its empirical formula (atomic mass of Fe = 56 μ)
3. Calculate the number of atoms in each of the following (i) 52 moles of Ar (ii) 52 u of He (iii) 52 g of He.
4. What will be the molarity of a solution, which contains 5.85 g of NaCl(s) per 500 mL?
5. If the concentration of glugose ($C_6H_{12}O_6$) in blood is 0.9 g L^{-1}, what will be the molarity of glugose blood?
6. Calculate the percentage of N in NH_3 molecule.
7. What will be the mass of one ^{12}C atom in g?
8. Caluculate the mass percentage of C in C_2H_4.
9. Calculate the mass per cent of different elements present in sodium sulphate, (Na_2SO_4).
10. Calculate the molecular mass of glucose ($C_6H_{12}O_6$) molecule.