

Model Question paper
Chemical Kinetics - I 1

11th Standard

Chemistry

Reg.No. :

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I. Answer all the questions.

II. Use blue pen only.

Time : 00:45:00 Hrs

Total Marks : 30

Part - A

2 x 1 = 2

- 1) $\text{mol.dm}^{-3}\text{sec}^{-1}$ is the unit of
(a) rate (b) rate constant (c) order (d) active mass
- 2) The elementary step with slow rate represents
(a) rate determining step (b) maximum rate step (c) third order rate (d) overall order

Part - B

1 x 2 = 2

- 3) Define half life period.

Part - C

2 x 3 = 6

- 4) What is a rate determining step?
- 5) List the factors on which an order of the reaction depend.

Part - D

4 x 5 = 20

- 6) Compare and contrast the terms, order and molecularity of a reaction.
- 7) Describe the factors on which the rate of a reaction depends.
- 8) What is a pseudo order reactions? How do you experimentally determine the pseudo first order rate constant of acid hydrolysis ester reaction?
- 9) Discuss the rate of the reaction



