

Model Question paper
Periodic Classification - I 1

11th Standard

Chemistry

Reg.No. :

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I. Answer all the questions.

II. Use blue pen only.

Time : 01:30:00 Hrs

Total Marks : 50

10 x 1 = 10

Part - A

- 1) The elements with atomic numbers 31 belongs to
(a) d-block (b) f-block (c) p-block (d) s-block
- 2) The most electronegative element of the periodic table is
(a) Iodine (b) Flourine (c) Chlorine (d) Oxygen
- 3) Which of the following forms stable gaseous negative ion.
(a) F (b) Cl (c) Br (d) I
- 4) The elements having highest ionisation energies within their periods are called:
(a) Halogens (b) Noble gases (c) Alkali metals (d) Transition elements
- 5) A property which progressively increases down a group in the periodic table is:
(a) Ionization enthalpy (b) Electronegativity (c) Electron gain enthalpy (d) Strength as a reducing agent
- 6) Elements whose atoms have their s and p-sub-levels complete are the:
(a) Normal elements (b) Transition elements (c) Halogens (d) Intergases
- 7) The law of octaves was stated by:
(a) Doberiner (b) Mendeleev (c) Mosely (d) Newland
- 8) Which of the following property decreases down a group:
(a) Ionization enthalpy (b) Atomic radii (c) Valency (d) All the above
- 9) Which of the following has the lowest melting point?
(a) CsCl (b) RbCl (c) KCl (d) NaCl
- 10) Excluding hydrogen and helium, the smallest element in the periodic table is:
(a) lithium (b) Oxygen (c) Fluorine (d) Chlorine

Part - B

5 x 2 = 10

- 11) Arrange F, Cl, Br and I in the order of increasing electronic gain enthalpy
- 12) Write electronic configurations for the elements of atomic numbers 6 and 14 and from this find out of which group in the periodic table each element belongs.
- 13) Which of the following electronic configuration has the lowest ionization enthalpy? $1s^2, 2s^2, 2p^6$
- 14) Which of the following electronic configuration has the lowest ionization enthalpy?
 $1s^2, 2s^2, sp^6$
- 15) State Modern Periodic Law

Part - C

5 x 3 = 15

- 16) Name any two transition elements and any two inner transition elements.
- 17) Why do elements in the same group have generally similar properties?
- 18) Arrange the order of increasing atomic volumes in
(a) Li, Na and K
- 19) Arrange the order of increasing atomic volumes in
(a) C, N and O
- 20) Name the different blocks of elements in periodic table. Give the general electronic configuration of each block.

Part - D

3 x 5 = 15

- 21) Which of the following pairs of elements would have higher electron gain enthalpy?
(a) N or O
- 22) Define transition elements. Name the different transition series
- 23) What do you mean by the term electron gain enthalpy? How does electron gain enthalpy change along a period and in a group?
