

Model Question paper
Group 2 s-Block elements 1
11th Standard

Chemistry

Reg.No. :

--	--	--	--	--	--

I. Answer all the questions.

II. Use blue pen only.

Time : 00:45:00 Hrs

Total Marks : 35

5 x 1 = 5

Part - A

- 1) Among the following, which is known as 'alkaline earth metal'.
(a) Sodium (b) Calcium (c) Lithium (d) Potassium
- 2) Alkaline earth metals are
(a) monovalent (b) trivalent (c) divalent (d) zerovalent
- 3) Among alkaline earth metals _____ is having the highest ionization energy.
(a) Beryllium (b) magnesium (c) Calcium (d) Barium
- 4) The colour given by barium in flame is
(a) Brick red (b) Apple Green (c) Red (d) Blue
- 5) The third most abundant dissolved ion in the ocean is
(a) Beryllium (b) Barium (c) Calcium (d) Magnesium

Part - B

2 x 2 = 4

- 6) Why the oxides of Group 2 metals have high melting points?
- 7) Why there is increase in the ionisation potential for forming M^{3+} ion for group 2 metals?

Part - C

2 x 3 = 6

- 8) Why do beryllium halides fume in air?
- 9) Why group 2 elements are harder than alkali metals?

Part - D

4 x 5 = 20

- 10) What are alkaline earth metals? Why are they called so?
- 11) How can you explain the anomalous behaviour of beryllium?
- 12) How does magnesium occur in nature? How is the metal extracted from its Ore?
- 13) Why the first ionization energy of alkaline earth metals higher than that of 1st group.
